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COLLEGE: ENGINEERING

MATRIC NO: 17/ENG04/021

DEPARTMENT: ELECT/ELECT

QUESTION 1

1. Step1- if the mass of the molecular ion is odd it contains at least one nitrogen N= 14 atoms 105-14=91

Step2- determine max NC’S

91/12 = 7.5 C7NH?

Sep3- add enough H’s to make up the rest of the madd

7×12=84

1×14=14

105-(84+14)=7

7H’S gives C7NH7

(2n+2-7)/2= 2(7.5)+2-7/2 =5.25

Step4- add an O atom

C7NH9→C6N0H3

(2(6.5) + 2−3)/2=5.5 ~6.

2 – Organic compounds are important because they serve as the basic form of all carbon bases for life on earth.

* Create energy production in biological life
* Causes atmospheric depletion and releases hydrocarbon energy
* Organic compounds have versatile bonding patterns and are part of all organisms
* Long carbon chain can be produced
* Will bond with many other elements
* Can form single, double and triple bonds
* A huge number of carbons is produced
* Organic compounds form stable bonds to other carbon

atoms- (catenation).

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| Homocyclic | Heterocyclic |
| They are cyclic compounds having atoms of the same element as ring members | They are cyclic compounds having atoms of different elements as ring members including carbon atoms |
| Ring contains atom of the same element | Ring contains atoms of different elements |
| Contains atoms of the same element bonded to each other containing a ring | Contains atoms of at least two different element bonded to each other forming a ring |
| Examples include: benzene, cyclohexane,toluene, cyclohexanol | Examples include: pyran, azocibe, thiocane etc. |

QUESTION 2

1. R.f of the first band = 2.4/12.2= 0.19=~ 0.2.

R.f of the second band= 5.6/12.2= 0.45=~ 0.5.

R.f of the third band= 8.9/12.2= 0.729=~ 0.73.

1. A- belongs to the family of the aldehyde, aromatic aldehyde and alpha hydroxyl ketone functional groups

B- belongs to the alkene or alkyne family.

1. Brandy’s test 2,4- Dinitrophenylhydrazine can be used to qualitatively detect the carbony functionality of a ketone or aldehyde functional group.

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| Organic compounds | Functional group | example |
| 1. Alkanes | RH | CH4- methane  C2H6- propane |
| 1. Alkenes | RR’  C=CR2R3CH3 | CH2=CH2- ethylene  CH2=CH2- propene |
| 1. Alkynes | RIC≡CR2 | HC≡ CH- acetylene  CH3 C ≡ CH HC≡ CH- propene |
| 1. Alcohols | ROH | CH3OH- methanol  C2H5OH- ethanol |
| 1. Alkyl halides | RX | CHCL3- chloroform  CH2CL2- dichloromethane |
| 1. Aldehyde | RCHO | CH3CHO- ethanal  CH2O- methanal |
| 1. Carboxylic acid | RCOOH | CH3COOH- ethanoic acid  HCOOH- formic acid |