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MATRIC NO: 17/ENG06/083

DEPARTMENT: MECHANICAL

QUESTION 1

* Fragment at m/z =105

Step1- if the mass of the molecular ion is odd it contains at least one nitrogen N= 14 atoms 105-14=91

Step2- determine max NC’S

 91/12 = 7.5 C7NH?

Sep3- add enough H’s to make up the rest of the madd

 7×12=84

 1×14=14

 105-(84+14)=7

 7H’S gives C7NH7

 (2n+2-7)/2= 2(7.5)+2-7/2 =5.25

Step4- add an O atom

 C7NH9→C6N0H3

 (2(6.5) + 2−3)/2=5.5 ~6.

* – Organic compounds are important because they serve as the basic form of all carbon bases for life on earth.
* Create energy production in biological life
* Causes atmospheric depletion and releases hydrocarbon energy
* Organic compounds have versatile bonding patterns and are part of all organisms
* Long carbon chain can be produced
* Will bond with many other elements
* Can form single, double and triple bonds
* A huge number of carbons is produced
* Organic compounds form stable bonds to other carbon

atoms- (catenation).

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| Homocyclic | Heterocyclic |
| They are cyclic compounds having atoms of the same element as ring members | They are cyclic compounds having atoms of different elements as ring members including carbon atoms |
| Ring contains atom of the same element  | Ring contains atoms of different elements  |
| Contains atoms of the same element bonded to each other containing a ring  | Contains atoms of at least two different element bonded to each other forming a ring  |
| Examples include: benzene, cyclohexane,toluene, cyclohexanol | Examples include: pyran, azocibe, thiocane etc.  |

QUESTION 2

* R.f of the first band = 2.4/12.2= 0.19=~ 0.2.

R.f of the second band= 5.6/12.2= 0.45=~ 0.5.

R.f of the third band= 8.9/12.2= 0.729=~ 0.73.

* A- belongs to the family of the aldehyde, aromatic aldehyde and alpha hydroxyl ketone functional groups

B- belongs to the alkene or alkyne family.

* Brandy’s test 2,4- Dinitrophenylhydrazine can be used to qualitatively detect the carbony functionality of a ketone or aldehyde functional group.
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| Organic compounds  | Functional group | example |
| * Alkanes
 | RH | CH4- methaneC2H6- propane |
| * Alkenes
 | RR’C=CR2R3CH3 | CH2=CH2- ethyleneCH2=CH2- propene |
| * Alkynes
 | RIC≡CR2 | HC≡ CH- acetyleneCH3 C ≡ CH HC≡ CH- propene |
| * Alcohols
 | ROH | CH3OH- methanolC2H5OH- ethanol |
| * Alkyl halides
 | RX | CHCL3- chloroformCH2CL2- dichloromethane |
| * Aldehyde
 | RCHO | CH3CHO- ethanalCH2O- methanal |
| * Carboxylic acid
 | RCOOH | CH3COOH- ethanoic acidHCOOH- formic acid |