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DEPARTMENT: MEDICINE AND SURGERY

MATRICULATION NUMBER: 19/MHS01/089

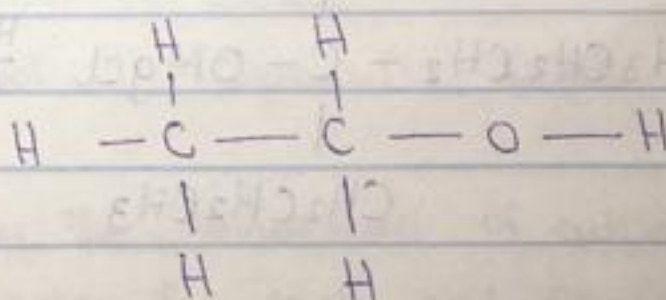
1. Discuss the two major classification of alkanols. Give two examples each for each class.

A. Classification based on the number of hydrogen atoms attached to the Carbon atom containing the hydroxyl group:

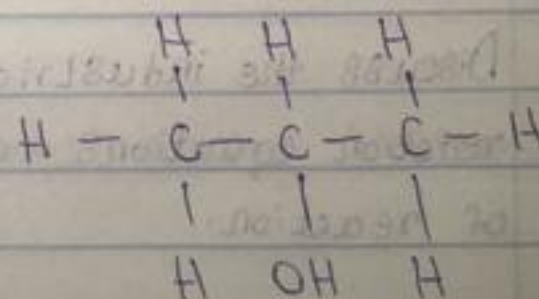
If the numbers of hydrogen atoms attached to the Carbon atom bearing the hydroxyl group are three or two, it is called a "Primary alkanol" ( $1^\circ$ ) [In a Primary alkanol, the hydroxyl group is attached to a Primary (or terminal) Carbon atom in the molecule, it is characterized by  $\text{CH}_2\text{OH}$ ]. If it is one hydrogen atom attached to the Carbon atom bearing the hydroxyl group, it is called "Secondary alkanol" ( $2^\circ$ ) [In a Secondary alkanol, the hydroxyl group is attached to a Secondary Carbon atom. It is characterized by  $>\text{CHOH}$ ] and if no hydrogen atom is attached to the Carbon atom bearing the hydroxyl group, it is called a "tertiary alkanol" ( $3^\circ$ ) [In a tertiary alkanol, the  $-\text{OH}$  is on a tertiary Carbon. It is characterized by  $>\text{C}-\text{OH}$ ].

Examples:

$\text{CH}_3\text{CH}_2\text{OH}$  - Ethanol ( $1^\circ$ )



$\text{CH}_3\text{OH}(\text{OH})\text{CH}_3$  - Propan-2-ol ( $2^\circ$ )



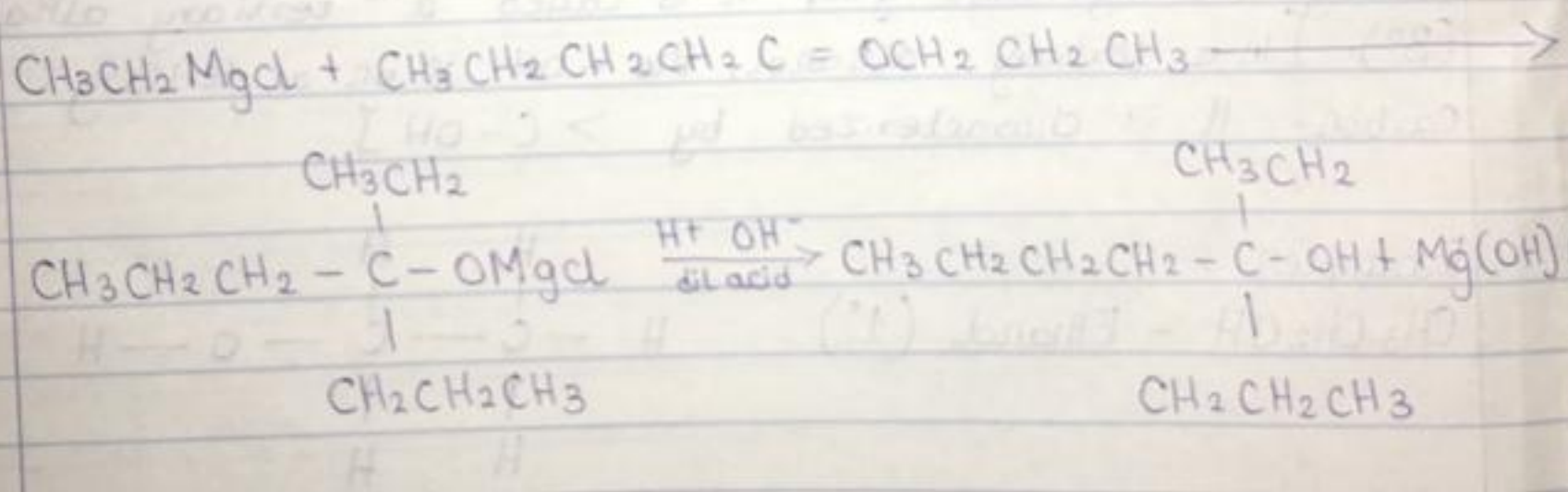
B classification based on the number of hydroxyl groups they have. Monohydric alkanols have only one hydroxyl group per molecule present in the alkanol structure. Dihydric alkanols also called Glycols have two hydroxyl groups present in the alkanol structure while Trihydric alkanols or triols have three hydroxyl groups present in the structure of the alkanol. Polyhydric alkanols or polyols have more than three hydroxyl groups.

Example

- i.  $\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$  - Propanol (monohydric alkanol)
- ii.  $\text{CH}_3\text{CH}(\text{OH})\text{CH}_2\text{CH}(\text{OH})\text{CH}_2\text{CH}_3$  - Hexane - 2,4 - diol (Dihydric alkanol)

2. In the Grignard synthesis of alkanols, react a named Grignard reagent with  $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{C}=\text{OCH}_2\text{CH}_2\text{CH}_3$ . Show the reaction steps.

Grignard reagent:  $\text{CH}_3\text{CH}_2\text{MgCl}$  (ethylmagnesium chloride).

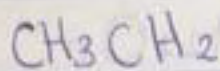
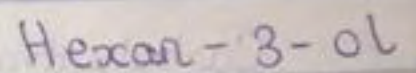
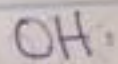
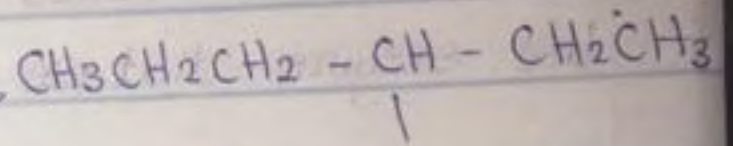
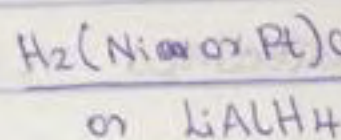
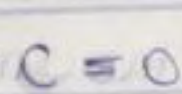
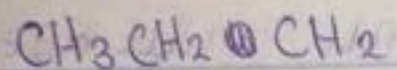


3. Discuss the industrial manufacture of ethanol showing all reaction equations and necessary enzymes and temperature of reaction.

Carbohydrates such as Starch are major groups of natural compounds that can be made to yield ethanol by the biological process of fermentation. The biological catalyst

The reaction above shows the reduction of an alkanal (eg. Butanal) to an alkanol (eg. butanol)

## ALKANONE

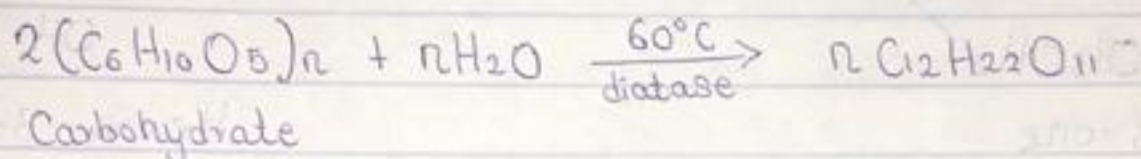


Hexan-3-one

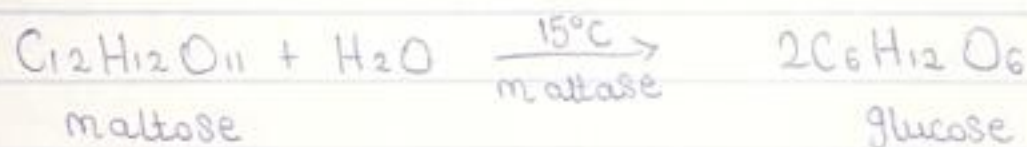
The reaction above shows the reduction of an alkanone (eg. Hexan-3-one) to an alkanol (Hexan-3-ol)

enzymes found in yeast breakdown the carbohydrate molecules into ethanol to give a yield of 95%.

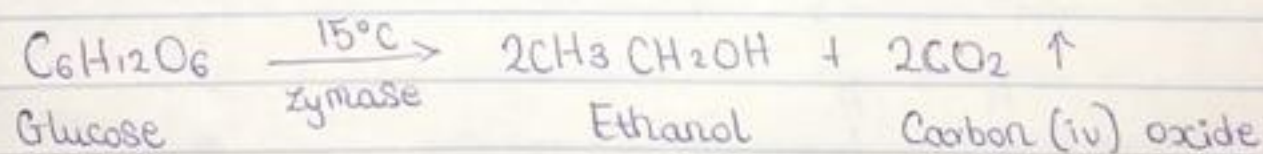
STEP 1: The Starch Containing materials include molasses, Potatoes, Cereals, rice and on warming with malt to  $60^{\circ}\text{C}$  for a specific period of time are converted into maltose by the enzyme diastase contained in the malt.



STEP 2: The maltose is broken down into glucose on addition of Yeast which contains the enzyme maltase and at a temperature of  $15^{\circ}\text{C}$



STEP 3: The glucose at constant temperature of  $15^{\circ}\text{C}$  is then converted into alcohol by the enzyme zymase also contained in yeast.



4. Determine the Product obtained by the reduction of alkanone and alkanal. Use a Specific example for each and Show the equation of reaction.

The reduction of alkanone and alkanal using MERWEIN-PONNDORF REACTION

## ALKANAL

