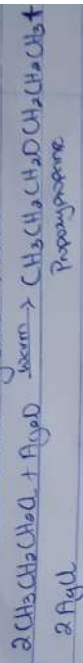


3. From Halobutanes and dry silver(I) oxide



Propoxypropane

4. Uses of Ethylene Oxide

1. Ethylene oxide is used as an intermediate in the hydrolytic manufacture of ethylene glycol.
2. Ethylene oxide is used in the preparation of nonionic emulsifying agents, plastics, plasticizers and several synthetic textiles.
3. Ethylene oxide is used as a gaseous sterilizing agent.

3. Density: Most of the simple ethers are less dense than water, although the density increases with increasing relative molecular mass and some of the aromatic ethers are in fact denser than water.

4. Boiling Point: Low molecular mass ethers have a lower boiling point than the four carbon alcohols, the reverse is true. The boiling point of ethers tend to approximate those of hydrocarbons of same relative molecular mass from which it can be concluded that the molecules are not associated in the liquid phase as there are no suitably available hydrogen for association through hydrogen bonds.

5. Reactivity: Ethers are inert at moderate temperature. Their inertness at moderate temperatures leads to their wide use as reaction media. Simple ethers are not found commonly in nature.

3. Methods of Preparing ethers

1. Controlled catalytic hydration of olefins



2-Isopropoxypropane

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CHM 102

Assignment III

- i.: CH_3OCH_3 - Methoxymethane
- ii.: $\text{CH}_3\text{CH}_2\text{OCH}_2\text{CH}_3$ - Ethoxyethane
- iii.: $(\text{CH}_3\text{CH}_2\text{CH}_2\text{O})_2\text{O}$ - Dibutyl ether
- iv.: $\text{CH}_3\text{CH}_2\text{OCH}_3$ - Methoxyethane
- v.: $\text{CH}_3\text{CH}_2\text{CH}_2\text{OCH}_2\text{CH}_3$ - Ethoxypropane

2. Properties of Ethers.

1. Physical States: At room temperature, ethers are colourless, neutral liquids with pleasant odours. The lower aliphatic ethers are highly flammable gases or volatile liquids.

2. Solubility: Ethers are less soluble in water than are the corresponding alcohols. Lower molecular weight ethers such as methoxymethane and methoxyethane are fairly soluble in water since the molecule are able to form hydrogen bonds with the water molecules but as the hydrocarbon content of the molecules increases, there is a rapid decline in solubility. They are miscible with most organic solvents.