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CIVIL ENGINEERING

11/ENIG13/22

1) CH_3OCH_3 - methoxymethane

$\text{CH}_3\text{CH}_2\text{OCH}_2\text{CH}_3$ - ethoxyethane

$\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{O}$ - butoxymethane

$\text{CH}_3\text{CH}_2\text{OCH}_2$ - methoxyethane

$\text{CH}_3\text{CH}_2\text{CH}_2\text{OCH}_2\text{CH}_3$ - ethoxypropane

2) Properties of ethers

Physical properties

1) Density: Most ethers are less dense than water but the density increases with increasing relative molecular mass.

2) Solubility: Ether molecules are miscible with most organic solvents.

3) Boiling point: Their boiling point is lower than that of alcohols.

4) Physical state: There are colorless, neutral liquids with pleasant odours at room temperature.

3) Three uses of ethylene oxide are

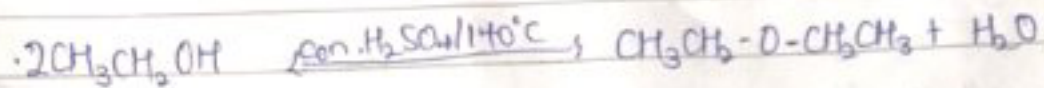
1) It is used as a gaseous sterilising agent.

2) It is used in the manufacture of detergents, fumigants etc.

3) It is used in the production of plastics, and synthetic textiles.

Methods Preparation of ethers

i) Partial dehydration of alcohols: This method includes acid which is tetraoxosulphate (vi) acid being heated. It includes acid catalysed and catalytic dehydration.



ii) From Alkyl halides: This Ethers can be prepared by heating haloalkanes with dry silver oxide.

