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**DEPARTMENT: DENTISTRY**

**COURSE CODE: CHM 102**

**MATRIC NUMBER: 19/MHS09/012**

1. **Naming of Carboxylic acids**
2. HCOOH – Methanoic Acid
3. HOOCCH2CH2CH2COOH – Pentan-1,5- dioic acid
4. CH3CH2CH2COOH- Butanoic acid
5. HO2C-CO2H- Ethanedioic acid
6. CH3(CH2)4COOH- Hexanoic acid
7. CH3CH=CHCH2CH2COOH- Hex-4-eneoic acid

**2. Physical properties**

**Physical appearances**

All simple aliphatic carboxylic acids up to C10 are liquids at room temperature. Most other carboxylic acids are solid at room temperature although anhydrous carboxylic acid (acetic acid) also known as glacial ethanoic acid freezes to an ice-like solid below the room temperature.

**Boiling points**

Boiling point increases with increasing relative molecular mass. Aromatic carboxylic acids are crystalline solids and have higher melting points than their aliphatic counterparts of comparable relative molecular mass.

**Solubility**

Lower molecular mass carboxylic acids with up to four carbon atoms in their molecules are soluble in water; this largely due to their ability to form hydrogen bonds with water molecules. The water solubility of the acids decreases as the relative molecular mass increases because the structure becomes relatively more hydrocarbon in nature and hence covalent. All carboxylic acids are soluble in organic solvents

1. **INDUSTRIAL PREPARATIONS OF CARBOXYLIC ACIDS**
2. **From Carbon(II) oxide**

Methanoic acid (formic acid) is manufactured by adding carbon(II)oxide under pressure to hot aqueous solution of sodium hydroxide. The free carboxylic acid is liberated by careful reaction with tetraoxosulphate (vi) acid (H2SO4)

CO--- **NaOH**-----🡪 HCOONa-----**H2SO4**--🡪 HCOOH + NaHSO4

1. **From ethanol**

Ethanoic acid is obtained commercially by the liquid phase air-oxidation of 5% solution of ethanal to ethanoic acid using manganite (II) ethanoate catalyst. Ethanal itself is obtained from ethylene

HC CH---- dil. H2SO4/HgSO4-🡪 CH3CHO-----O2/ (CH3COO)2Mn---🡪 CH3COOH

**4.** **SYNTHETIC PREPARATIONS**

 **i. Oxidation of primary alcohols and aldehydes**

Oxidation of primary alcohols and aldehydes can be used to prepare carboxylic acids using the usual oxidizing agents (i.e K2Cr2O7 or KMnO4) in acidic solution

RCH2OH------[O], excess acid/KMnO4----🡪 RCHO------ [O] ------🡪 RCOOH

1. **Carbonation of Grignard reagent**

Aliphatic carboxylic acids are obtained by bubbling carbon (IV) oxide into the Grignard reagent and then hydrolyzed with dilute acid

RMgBr + CO2-------- (C2H5)2O---🡪 RCOOMg------Br/H2O/ dil. acid----🡪 RCOOH + MgBrOH

R may be 1o, 2o , 3o aliphatic alkyl or aryl radical

In the preparation of benzoic acid, the reagent is added to solid carbon (IV) oxide (dry ice) which also serves as coolant to the reaction mixture

C6H5MgBr + CO2------ (C2H5)2O---🡪 C6H5COOMgBr------H2O/H+----🡪 C6H5COOH + MgBrOH

1. **Hydrolysis of nitriles (cyanides) or esters**

RCN + 2H2O------H+----🡪 RCOOH + NH4+

(R=alkyl or aryl radical)

RCOOR’ ------------H2O/H+ reflux------------🡪 RCOOH + R’OH

C6H5CH2CN + 2H2O----------H+--------🡪 C6H5CH2COOH + NH4+

 CH3CH2COOCH3 ------------H2O/H+ reflux------------🡪CH3CH2CO

**5. i. Reduction to primary alcohol**

4RCOOH + 3LiAlH4--------(C2H5)2O------🡪(RCH2O)4AlLi + 2LiAlO2 + 4H2

 4H2O

4RCH2OH + Al(OH)3 + LiOH

CH3CH2CH2COOH--------LiAlH4------🡪 CH3CH2CH2CH2OH

Butanoic acid Butanol

**ii. Decarboxylation**

Thermal decarboxylation

CH3CH2CH2COONa + NaOH ----------Fuse------🡪 CH3CH2CH3 + Na2CO3

Kolbe synthesis

2CH3CH2COONa + 2H2O ------electrolysis/aq. CH2OH----🡪 CH3(CH2)CH3 + CO2(anode) + 2NaOH + H(cathode)

iii. **Esterification**

CH3CH2CH2COOH + CH3CH2CH2OH <----H+----🡪CH3CH2CH2COOCH2CH2CH3 + H2O