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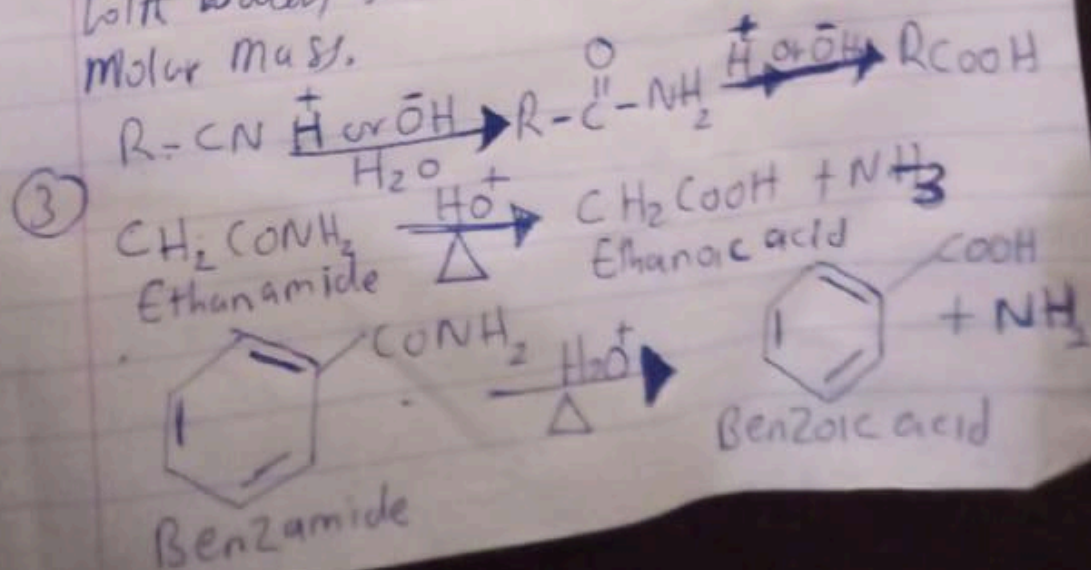
Department: Pharmacy

Matric No: 19/MHS11/061

Course: Chem

- ①  $\text{HCOOH}$  - Formic acid:  $\text{HCO}_2\text{H}$
- $\text{HOOCCH}_2\text{CH}_2\text{CH}_2\text{COOH}$  - ~~Propanoic~~ 1,3 Propanedioic
  - $\text{CH}_3\text{CH}_2\text{CH}_2\text{COOH}$  - This is butanoic acid
  - $\text{HO}_2\text{C}-\text{CO}_2\text{H}$  - dicarboxylic
  - $\text{CH}_3(\text{C}_2\text{H}_5)_4\text{COOH}$  - Butyric acid

② Carboxylic acid have a high boiling point compared to other substances of comparable to other substances of comparable mass. as for the boiling point, it increases with its molar mass, because carboxylic acid have one to four carbon atom miscible with water, while solubility decreases with molar mass.



⑤ All acid derivatives can be hydrolyzed to yield carboxylic acid, it requires range from mild to severe, depending on the compound involved

