

NWOKORIE PASCAL CHINAMDI

MECHANICAL ENGINEERING

19/ENGG051042

CHM 102 ASSIGNMENT [CARBOXYLIC ACID]

1) $\text{HCOOH} \rightarrow$ Methanoic acid

- $\text{HOOCCH}_2\text{CH}_2\text{CH}_2\text{COOH} \rightarrow$ Pentan-1,5-dioic acid

- $\text{CH}_3\text{CH}_2\text{CH}_2\text{COOH} \rightarrow$ Butanoic acid

- $\text{HO}_2\text{C}-\text{CO}_2\text{H} \rightarrow$ Ethanedioic acid.

- $\text{CH}_3(\text{CH}_2)_4\text{COOH} \rightarrow$ Hexanoic acid

- $\text{CH}_3\text{CH}=\text{CHCH}_2\text{CH}_2\text{COOH} \rightarrow$ ~~Penten~~ 4-Hexenoic acid

2) Physical Properties

- Physical appearance:

All simple aliphatic carboxylic acids up to C_{10} are liquids at room temperature. Most other carboxylic acids are solid at room temperature although anhydrous carboxylic acid (acetic acids) also known as glacial ethanoic acid freezes to an ice-like solid below the room temperature.

- Boiling Point:

~~Boil.~~ Boiling point increases with relative molecular mass. Aromatic carboxylic acids are crystalline solids and have higher melting points than their aliphatic counterparts of comparable relative molecular mass.

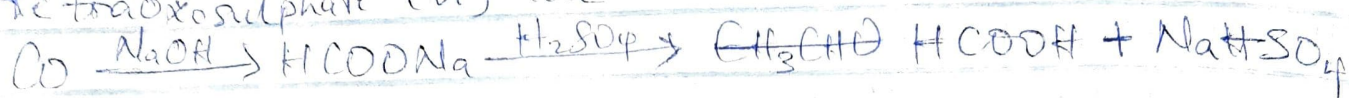
- Solubility:

Lower molecular mass carboxylic acids with up to four carbon atoms in their molecules are soluble in water; this is largely due to their ability to form hydrogen bonds with water molecules. The water solubility of the acids decreases as the relative molecular mass increases because the structure becomes relatively more hydrocarbon in nature and hence covalent. All carboxylic acids are soluble in organic solvents.

3) Industrial preparations of carboxylic acids.

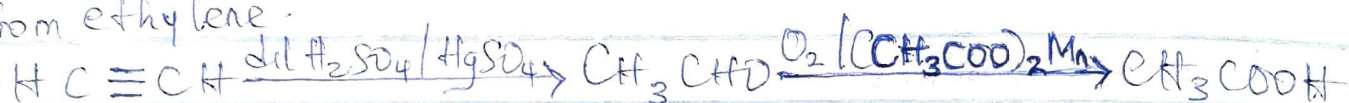
- ~~From~~ From Carbon (II) oxide:

Methanoic acid (Formic acid) is manufactured by adding Carbon (iv) oxide under pressure to hot aqueous solution of Sodium hydroxide. The free carboxylic acid is liberated by careful reaction with tetraoxosulphate (vi) acid.



- From Ethanol:

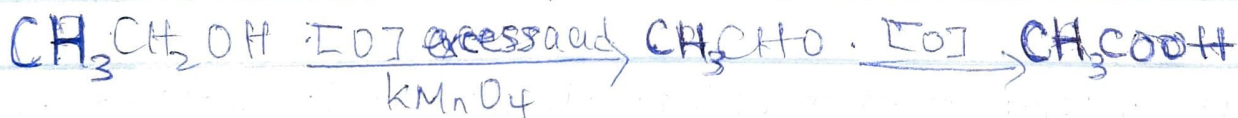
Ethanoic acid is obtained commercially by the liquid phase air-oxidation of 5% solution of ethanal to ethanoic acid using ~~mang~~ manganite (ii) ethanoate catalyst. Ethanal itself is obtained from ethylene.



4) ~~Synth~~ Synthetic Preparation of Carboxylic acid.

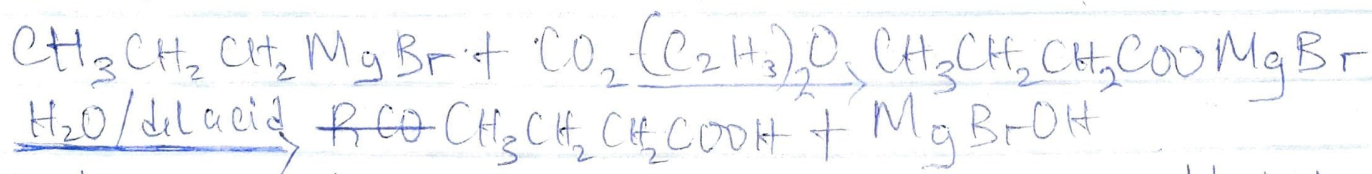
- Oxidation of Primary alcohols and aldehydes:

Oxidation of primary alcohols and aldehydes can be used to prepare carboxylic acids using the usual oxidizing agents (i.e. $\text{K}_2\text{Cr}_2\text{O}_7$ or KMnO_4) in acidic solution.

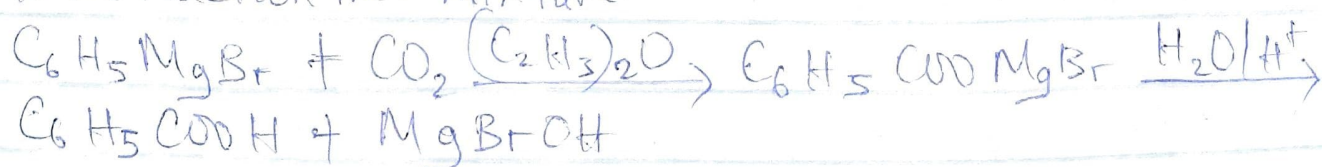


- Carbonation of Grignard Reagent:

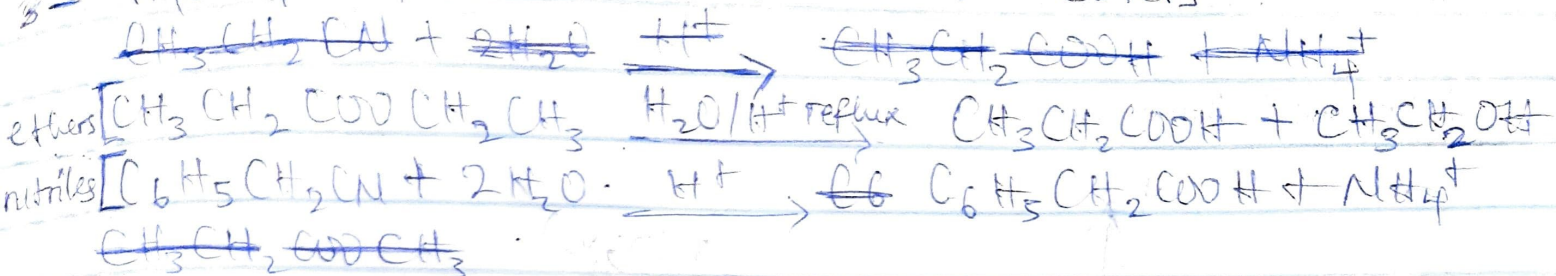
Aliphatic carboxylic acids are obtained by bubbling Carbon (iv) oxide into the Grignard reagent and then hydrolyzed with dilute acid.



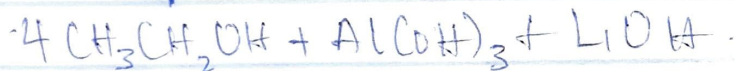
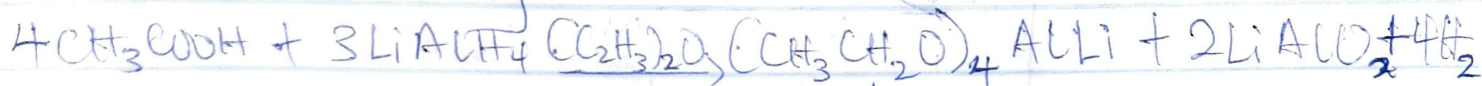
In the preparation of benzoic acid; the reagent is added to solid carbon (iv) oxide (dry ice) which also serves as coolant to the reaction mixture.



3 - Hydrolysis of Nitriles (or amides) or esters



5) - Reduction to Primary alcohol

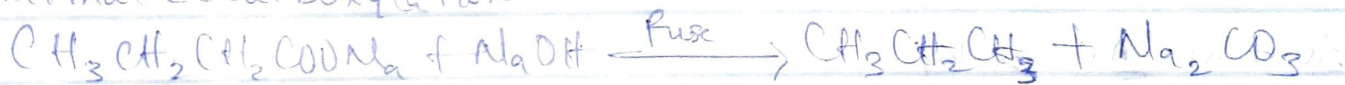


Butanoic acid

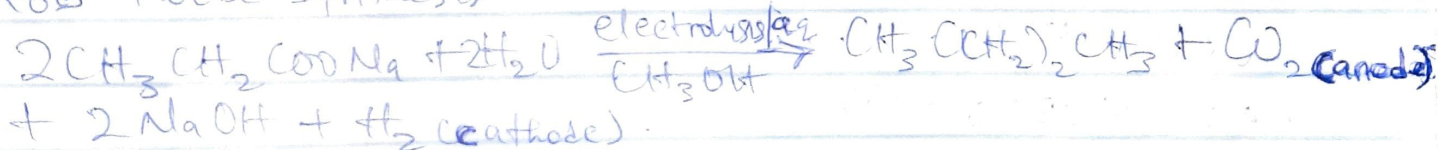
Butanol

2 - Decarboxylation

Thermal decarboxylation



Kolb Kolbe Synthesis



- Esterification

