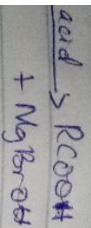




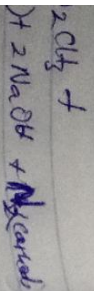
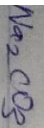
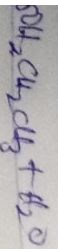
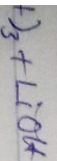
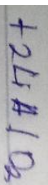
Hydros can be
oxidizing agents

giving CO_2 in the
acid.



added to stirred
reaction mixture.

OH



Name: - OJUNIBANWO DUNFACISTO SIKEMI

MATRIC NO: - 191MHSOL/298

DEPARTMENT: - MBBS

1.) HCOOH - methanoic acid

• HOOCC_2H_5 - Butanoic acid

• $\text{CH}_3\text{CH}_2\text{CH}_2\text{COOH}$ - Pentanoic acid

• $\text{H}_2\text{C}(\text{C}_2\text{H}_5)_2$ - Hexanoic acid

• $\text{CH}_3(\text{CH}_2)_4\text{COOH}$ - Hexanoic acid

• $\text{CH}_3(\text{CH}_2)_5\text{COOH}$ - Hex-decanoic acid

2.) Physical appearances: - All simple aliphatic carboxylic acids up to C_6 are liquids at room temperature. Most other carboxylic acids are solid at room temperature although anhydrous carboxylic acids are solid at room temperature although anhydrous carboxylic acids are solid at room temperature.

(Carboxylic acid) Freezes to a ice-like solid below the room temperature.

ii.) Boiling point: - It increases with increasing relative molecular mass.

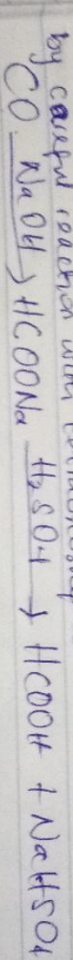
Aromatic carboxylic acids are crystalline solids and have higher melting points than their aliphatic counterparts of comparable relative molecular mass.

iii.) Solubility: - Carboxylic acids of lower molecular mass with up to four carbon atoms in their molecules are soluble in water due to their ability to form hydrogen bonds with water molecule. Water solubility of acid decreases with increase in relative molecular mass because the structure becomes more hydrocarbon in nature.

All carboxylic acids are soluble in organic solvent.

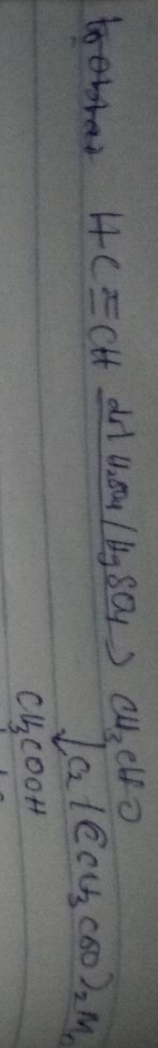
3.) - From Carbon(II)oxide

Methanoic acid is produced by adding CO under pressure to hot aqueous solution of NaOH . The free carboxylic acid is liberated by careful reaction with tetraoxosulphate (VI) acid.

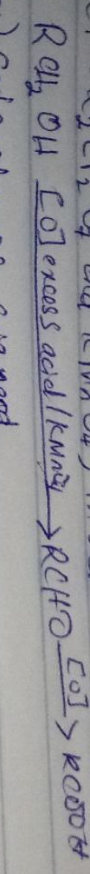


- From ethanol

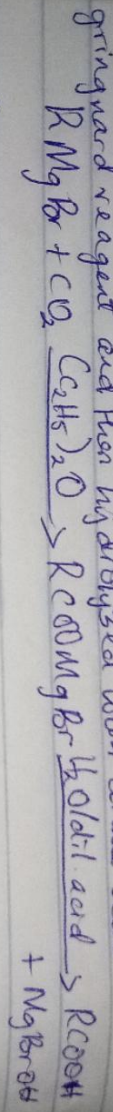
Ethanoic acid is obtained commercially by the liquid phase air-oxidation of C_2H_5 solution of ethanol to ethanoic acid using magnesium(II) ethanoate catalyst. Ethanol itself is obtained from ethylene.



4) Oxidation of primary alcohols and aldehydes
 Oxidation of primary alcohols and aldehydes can be used to prepare carboxylic acid using the usual oxidizing agent (i.e. $\text{K}_2\text{Cr}_2\text{O}_7$ and KMnO_4) in acidic solution

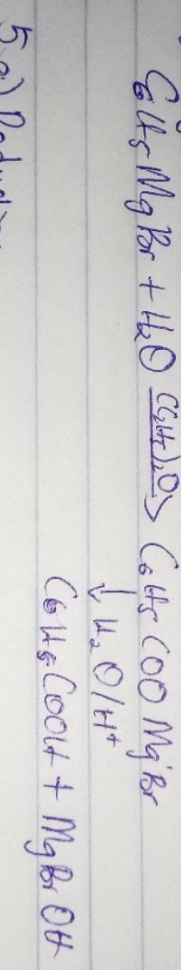


b) Carbonation of Grignard
 Aliphatic carboxylic acids are obtained by bubbling CO_2 into the Grignard reagent and then hydrolyzed with dilute acid.

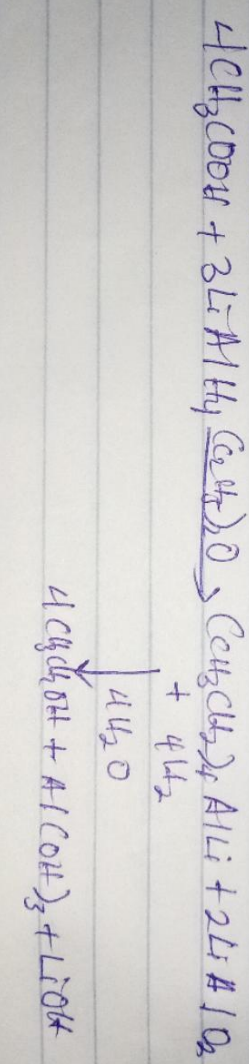


R may be 1°, 2°, 3° aliphatic alkyl radical.

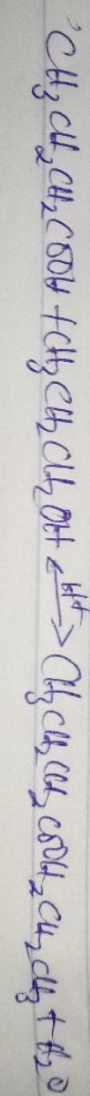
In preparation of benzoic acid, the reagent is added to solid CaC_2 (dry ice) which also serves as coolant to the reaction mixture.



5) Reduction



b) Esterification



a) Decarboxylation

Thermal decarboxylation

