

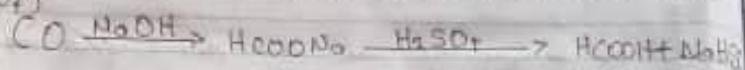
this largely due to their ability to form hydrogen bonds with water molecules. The water solubility of the acids decreases as the relative molecular mass increases because the structure becomes relatively more hydrocarbon in nature and hence. Covalent. All Carboxylic acids are soluble in organic solvents.

3. Write two industrial preparations of Carboxylic acids.

Answer

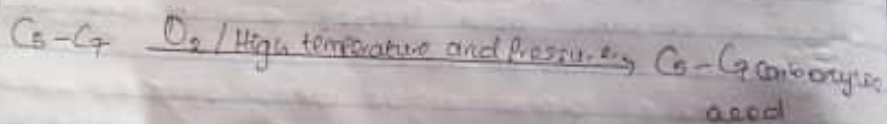
i) From Carbon Dioxide

Malonic acid (Ethane acid) is manufactured by reacting Carbon Dioxide under pressure to hot aqueous solution of Sodium hydroxide. The free Carboxylic acid is liberated by careful reaction with tetraoxosulphate (VI) acid (H_2SO_4)



ii) From Petroleum

Liquid phase air oxidation of C_5-C_7 alkanes, obtainable from petroleum at high temperature and pressure will give C_5-C_7 carboxylic acids with methanoic, propanoic and butanedioic acid as by-products.

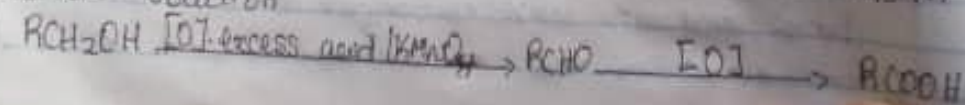


4) Write equations and brief explanation discuss the synthetic preparation of carboxylic acid.

Answer

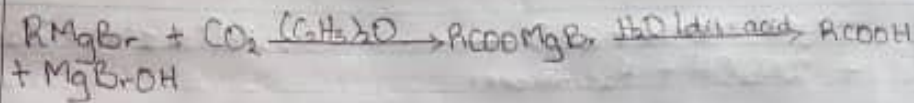
i) Oxidation of primary alcohols and aldehydes

Oxidation of primary alcohols and aldehydes can be used to prepare carboxylic acids using the usual oxidizing agents (i.e. $K_2Cr_2O_7$ or $KMnO_4$) in acidic solution

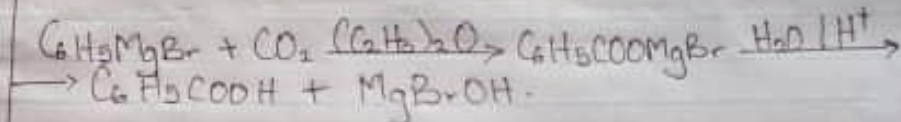


(ii) Carboxylation of Grignard reagent

Aliphatic carboxylic acids are obtained by bubbling carbon(IV)oxide into the Grignard reagent and then hydrolyzed with dilute acid.



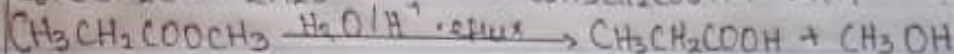
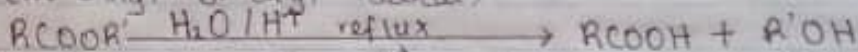
R may be 1°, 2°, 3° aliphatic alkyl or aryl radical
 In the preparation of benzoic acid, the reagent is added to solid carbon (IV)oxide (dry ice) which also serves as coolant to the reaction mixture.



(iii) Hydrolysis of nitriles (cyanides) or esters



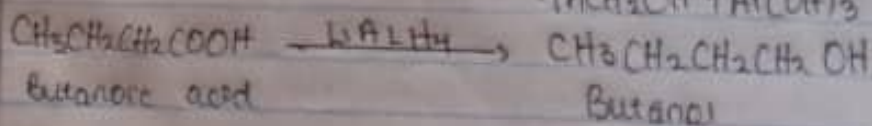
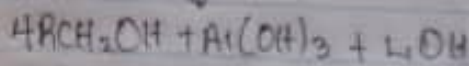
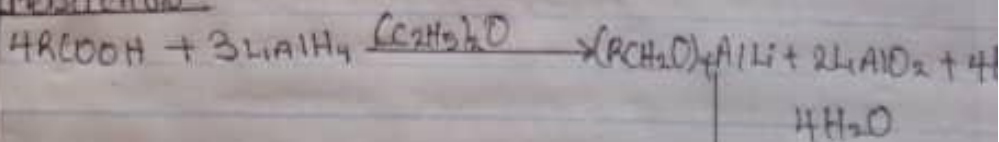
(R = alkyl or aryl radical)



5) With chemical equation only, outline the reduction, decarboxylation and esterification of carboxylic acid.

Answer

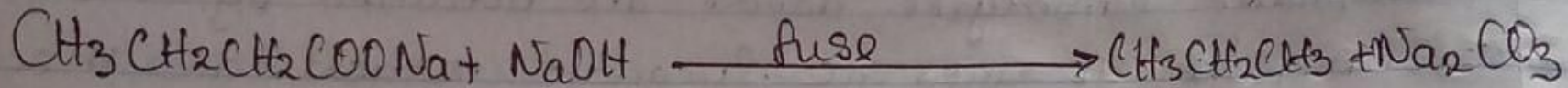
Reduction



butanoic acid

Butanol

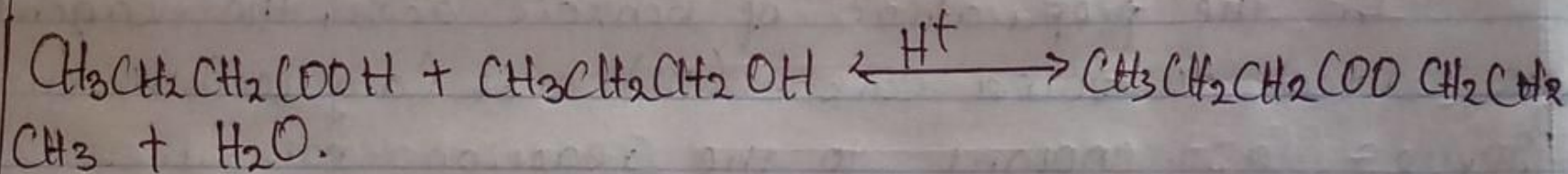
(i) Decarboxylation



Kolbe Synthesis



(ii) Esterification



Chemistry 102, Matric Number: 19/MHS01/056

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1) Give the IUPAC names of the following compounds

Answer

a) HCOOH \longrightarrow Methanoic acid.

b) $\text{HOOCCH}_2\text{CH}_2\text{CH}_2\text{COOH}$ \longrightarrow pentan-1, 5-diene acid

c) $\text{CH}_3\text{CH}_2\text{CH}_2\text{COOH}$ \longrightarrow Butanoic acid

d) $\text{HO}_2\text{C}-\text{CO}_2\text{H}$ \longrightarrow Ethanedioic acid.

e) $\text{CH}_3[\text{CH}_2]_4\text{COOH}$ \longrightarrow Hexanoic acid.

f) $\text{CH}_3\text{CH}=\text{CHCH}_2\text{CH}_2\text{COOH}$ \longrightarrow Hex-4-enoic acid.

2) Discuss Briefly the physical properties of carboxylic acids under the following headings.

i) PHYSICAL APPEARANCE

All simple aliphatic carboxylic acids up to C_{10} are liquids at room temperature. Most other carboxylic acids are solids at room temperature although anhydrous carboxylic acid (acetic acid) also known as glacial ethanoic acid freezes to an ice-like solid below the room temperature.

ii) Boiling points

Boiling points increases with increasing relative molecular mass. Aromatic carboxylic acids are crystalline solids and have higher melting points than their aliphatic counterparts of comparable relative molecular mass.

SOLUBILITY

Lower molecular mass carboxylic acid with up to four carbon atoms in their molecules are soluble in water.