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DEPARTMENT: NURSING (100 LEVEL)

COLLEGE: MEDICINE AND HEALTH SCIENCES

COURSE: CHM102

MATRIC NO: 19/MHS02/110

 ASSIGNMENT

1. The IUPAC names of the following compounds:

I. HCOOH- Methanoic acid

II. HOOCCH2CH2CH2COOH- Pentan-1, 5-dioic acid

III. CH3CH2CH2COOH- Butanoic acid

IV. HO2C-CO2H- Ethanedioic acid

V. CH3(CH2)4COOH- Hexanoic acid

VI. CH3CH=CHCH2CH2COOH- Hex-4-eneoic acid

2. Discuss briefly the physical properties of carboxylic acids:

 PHYSICAL APPEARANCES: All simple aliphatic carboxylic acids up to C10 are liquids at room temperature. Most other carboxylic acids are solid at room temperature although anhydrous carboxylic acid (acetic acid) also known as glacial ethanoic acid freezes to an ice-like solid below the room temperature.

BOILING POINTS: Boiling point increases with increasing relative molecular mass. Aromatic carboxylic acids are crystalline solids and have higher melting points than their aliphatic counterparts of comparable relative molecular mass.

SOLUBILITY: Lower molecular mass carboxylic acids with up to four carbon atoms in their molecules are soluble in water; this largely due to their ability to form hydrogen bonds with water molecules. The water solubility of the acids decreases as the relative molecular mass increase because the structure becomes relatively more hydrocarbon in nature and hence covalent. All carboxylic acids are soluble in organic solvents.

3. Write two industrial preparation of carboxylic acid:

I. From Carbon (II) oxide: Methanoic acid (formic acid) is manufactured by adding carbon (II) oxide under pressure to hot aqueous solution of sodium hydroxide. The free carboxylic acid is liberated by careful reaction with tetraoxosulphate (VI) acid (H2SO4)

 CO NaOH HCOONa H2SO4 HCOOH + NaHSO4

II. From ethanal: Ethanoic acid is obtained commercially by the liquid phase air-oxidation of 5% solution of ethanal to ethanoic acid using manganite (II) ethanoate catalyst. Ethanal itself is obtained from ethylene.

HC≡CH dil. H2SO4/HgSO4 CH3CHO O2/ (CH3COO)2 Mn CH3COOH

4. With equations and brief explanation discuss the synthetic preparation of carboxylic acid:

I. Oxidation of primary alcohols and aldehydes: Oxidation of primary alcohols and aldehydes can be used to prepare carboxylic acids using the usual oxidizing agents (i.e. K2Cr2O7 or KMnO4­) in acidic solution

RCH2OH [O], excess acid/ KMnO4 RCHO [O] RCOOH

II. Carbonation of Grignard reagent: Aliphatic carboxylic acids are obtained by bubbling carbon (IV) oxide into the Grignard reagent and then hydrolyzed with dilute acid

RMgBr + CO2 (C2H5)2O RCOOMgBr H2O / dil. acid RCOOH + MgBrOH

R may be 1 ͦ, 2 ͦ, 3 ͦ aliphatic alkyl or aryl radical

In the preparation of benzoic acid, the reagent is added to solid carbon (II) oxide (dry ice) which also serves as coolant to the reaction mixture

C6H5MgBr + CO2 (C2H5)2O C6H5COOMgBr H2O/ H+ C6H5COOH + MgBrOH

III. Hydrolysis of nitriles (cyanides) or esters

 RCN + 2H2O H+ RCOOH + NH4+

 (R= alkyl or aryl radical)

 RCOOR’ H2O/H+ reflux RCOOH + ROH

 C6H5CH2CN + 2H2O H+ C6H5CH2COOH + NH4+

 CH3CH2COOCH3 H2O/H+ reflux CH3CH2COOH + CH3OH

5. With chemical equation only, outline the reduction, decarboxylation and esterification of carboxylic acid:

I. REDUCTION

4RCOOH + 3LiALH4 (C2H5)2O (RCH2O)44H2O AILi + 2LiALO2 + 4H2 4H2O 4RCH2OH + AL(OH)3 + LiOH

CH3CH2CH2COOH LiALH4  CH3CH2CH2CH2OH

Butanoic acid Butanol

II. DECARBOXYLATION

CH3CH2CH2COONa + NaOH fuse CH3CH2CH3  + Na2CO3

Kolbe synthesis

2CH3CH2COONa + 2H2O electrolysis/aq.CH3OH CH3(CH2)2CH3 + CO2(anode) + 2NaOH + H2(cathode)

III. Esterification

CH3CH2CH2COOH + CH3CH2CH2OH H+ CH3CH2CH2COOCH2CH2CH3 + H2O