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1) HCOOH - Methanoic acid

$\text{HOOCCH}_2\text{CH}_2\text{CH}_2\text{COOH}$ - propane-1,2,3-tricarboxylic

$\text{CH}_3\text{CH}_2\text{CH}_2\text{COOH}$ - butanoic acid

$\text{HO}_2\text{CCO}_2\text{H}^+$ - Ethanedioic acid

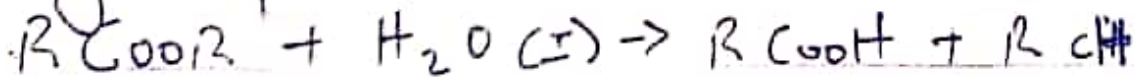
$\text{CH}_2\text{CC}(\text{H}_2)_4\text{COOH}$ - Hexanoic-1,6-dioic acid

2) physical appearance - They are solids with low melting point

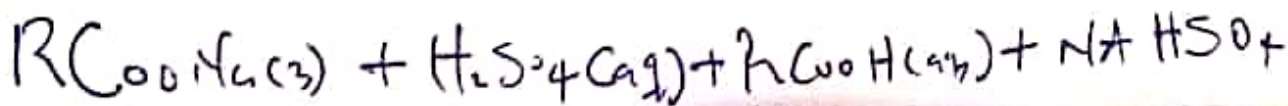
i) Boiling point - The simplest carboxylic acid boils at 101°C (214°F)

ii) Solubility - It is soluble in water

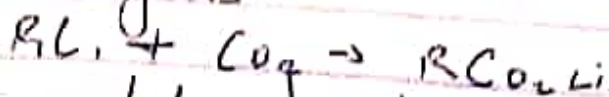
3) Hydrolysis of esters in acid medium



ii) Distillation of anhydrous sodium ethanoate with concentrated H_2SO_4



4) Carbonation of a Grignard reagent and organo lithium reagents



Base-catalyzed cleavage of non-enolized ketone especially aryl ketones

