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- Q Give the IUPAC names of the following compounds
- $\text{HCOOH} \rightarrow$ Methanoic acid
 - $\text{HOOCCH}_2\text{CH}_2\text{CH}_2\text{COOH} \rightarrow$ propanoic 5-dicar acid
 - $\text{CH}_3\text{CH}_2\text{CH}_2\text{COOH} \rightarrow$ Butanoic acid
 - $\text{HOOC}-\text{CO}_2\text{H} \rightarrow$ Ethanedioic acid
 - $\text{CH}_3(\text{CH}_2)_4\text{COOH} \rightarrow$ Hexanoic acid.
 - $\text{CH}_3\text{CH}=\text{CHCH}_2\text{CH}_2\text{COOH} \rightarrow$ Hex-4-enoic acid

e) Discuss Briefly the physical properties of carboxylic acid under the following headings.

i) Physical Appearance

All simple aliphatic carboxylic acids up to C_{10} are liquids at room temperature. Most other carboxylic acids are solids at room temperature although anhydrous carboxylic acid (acetic acid) also known as glacial ethanoic acid freezes to an ice like solid below the room temperature.

ii) Boiling points

Boiling points increases with increasing relative molecular mass aromatic carboxylic acids eg crystalline solids and have higher boiling points, than their aliphatic counterparts of comparable relative molecular mass.

iii) Solubility

Lower molecular mass carboxylic acid with up to four carbon atoms in their molecules are soluble in water, this is due to their ability to form hydrogen bonds with water molecules.

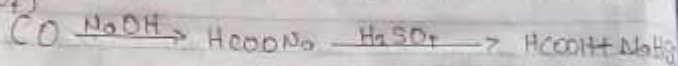
this, largely due to their tendency to form hydrogen bonds with water molecules. The water solubility of the acids decreases as the relative molecular mass increases because the structure becomes relatively more hydrocarbon in nature and hence, covalent. All carboxylic acids are soluble in organic solvents.

3. Write two industrial preparations of carboxylic acids.

Answer

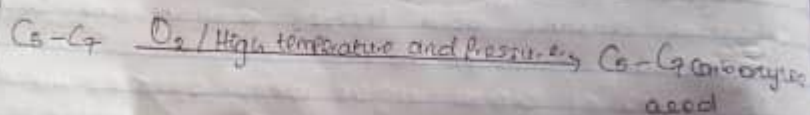
i) From Carbon Dioxide

Methanoic acid (formic acid) is manufactured by the carbon dioxide under pressure to hot aqueous solution of sodium hydroxide. The free carboxylic acid is then freed by carbon reaction with tetraoxosulphate (VI) acid (H_2SO_4)



ii) From Petroleum

Liquid phase air oxidation of C_5 - C_7 alkanes, obtainable from petroleum at high temperature and pressure will give C_5 - C_7 carboxylic acids with methanoic, propanoic and butanedioic acid as by-products.

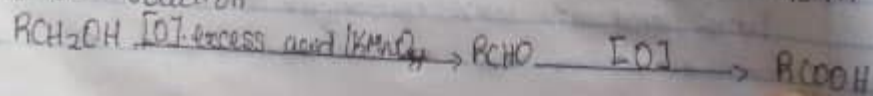


4) With equations and brief explanation discuss the synthetic preparation of carboxylic acid.

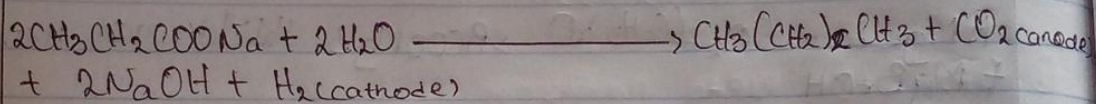
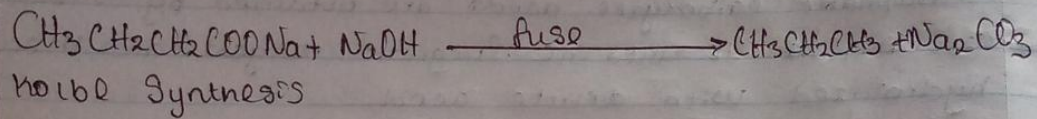
Answer

① Oxidation of primary alcohols and aldehydes

Oxidation of primary alcohols and aldehydes can be used to prepare carboxylic acids using the usual oxidizing agents (i.e. $\text{K}_2\text{Cr}_2\text{O}_7$ or KMnO_4) in acidic solution



(ii) Decarboxylation



(iii) Esterification

