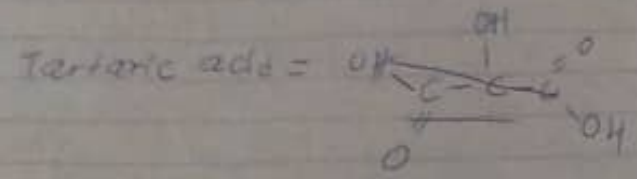
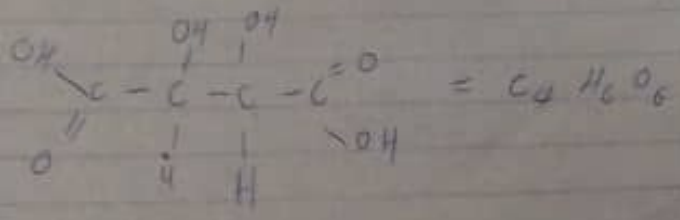
- Hydroxyl group (-OH)
- Alkanol (C=O, H)

2) Concentration ($mol\,dm^{-3}$) = Conc (g/cm^3)

$$[\alpha]^T = \frac{\alpha}{c \cdot l}$$

Tartaric acid =





$= C_4H_6O_6$

Molar mass = 150 g/mol

0.856 g = 10 cm^3

x g = 1000 cm^3

$$\frac{0.856 \times 1000}{10} = 85.6\ g/dm^3$$

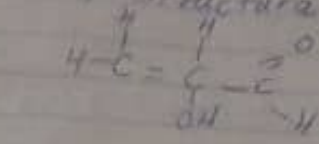
$$\text{Concentration in } g/cm^3 = \frac{\text{concentration (g/dm}^3\text{)}}{1000}$$

$$= \frac{85.6}{1000} = 0.0856\ g/cm^3$$

NAME: ALLAN STEPHEN NONSO
MATRIC: 191
Chemistry

i) $CH_2 = C(OH)HCHO$

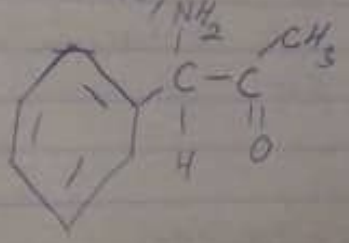
The structural formulae



Functional present are;
- Double bond chain = (Alkene)
- OH (hydroxyl group)
- $-C(=O)-$ (alkanal)

ii) $C_6H_5CH(NH_2)COCH_3$

Structure;



Functional present
- Phenyl group (C_6H_5) with double bonds
- Amine
- Alkenone/ketone ($C(=O)-R$)

iii) ~~(H)~~

$CH_3C = CHCH(OH)CHO$

Structure