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(1) Give the IUPAC names of the following compounds

- HCOOH (Methanoic acid)
- $\text{HOOC(CH}_2\text{)}_3\text{COOH}$ (pentan-1,5-dioic acid)
- $\text{CH}_3\text{CH}_2\text{CH}_2\text{COOH}$ (Butanoic acid)
- $\text{HO}_2\text{C}-\text{CO}_2\text{H}$ (Ethanedioic acid)
- $\text{CH}_3(\text{CH}_2)_4\text{COOH}$ (Hexanoic acid)
- $\text{CH}_3\text{CH}=\text{CHCH}_2\text{CH}_2\text{COOH}$ (Hex-4-enoic acid)

(2) Physical appearance: All simple aliphatic carboxylic acids up to C_{10} are liquid at room temperature. Most other carboxylic acids are solid at room temperature although anhydrous carboxylic acid (acetic acid) also known as glacial ethanoic acid ~~is~~ freezes to an ice-like solid below the room temperature.

b Boiling point: This increases with increasing relative molecular masses. Aromatic carboxylic acids are crystalline solids & have ~~higher~~ higher melting points than their aliphatic ~~comp~~ counterparts of comparable relative mass

c Solubility: lower molecular mass carboxylic acids with up to four carbon atoms in their molecules are soluble in water. This is largely due to their ability to form hydrogen bonds with water molecules

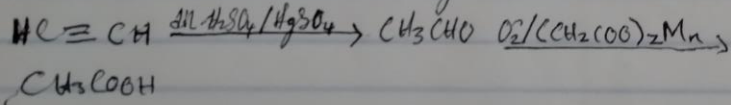
The water solubility of the acids decreases as the ~~solubility~~ relative molecular mass increases because the structure becomes

relatively more hydrocarbon in nature & hence Covalent
 All Carboxylic acids are soluble organic solvents.

(3) From petroleum L.P.G. Liquid phase air oxidation of
 $C_6 - C_7$ alkenes, obtainable from obtainable from
 petroleum at high temperature and pressure will give
 $C_5 - C_7$ Carboxylic acids with methanoic, propanoic
 & butanoic acids as by products
 $C_5 - C_7$ O_2 / high temperature & pressure, $C_5 - C_7$ Carboxylic
 acid

b From ethanal: ethanoic acid is obtained commercially by the
 liquid phase air-oxidation of 5% solution of ethanal
 to ethanoic acid using manganate(II) ethanoate catalyst.

Ethanal itself is obtained from ethylene



(4) Hydrolysis of nitriles (Cyanides) or esters

