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MATRIC NO:- 19/MHS 061001

DEPT:- MEDICAL LABORATORY SCIENCE

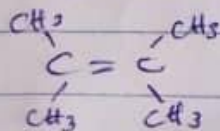
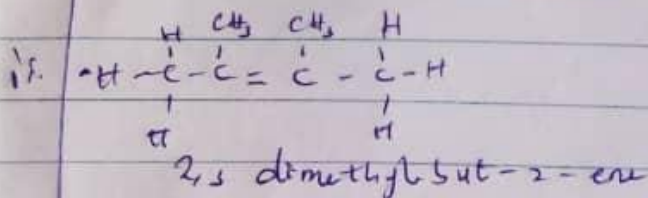
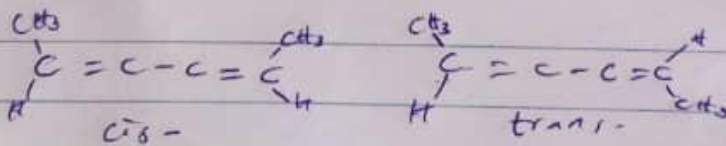
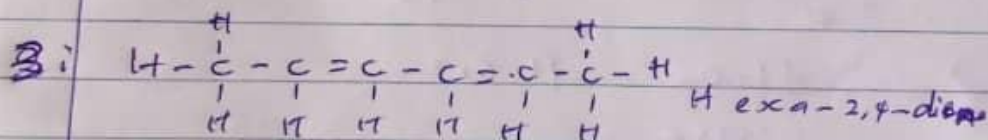
COURSE:- CHM 102

### Assignment

- (a) Formyl group (Aldehyde) group (CHO)  
(b) Hydroxyl group (OH)  
(c) Alkene group (double bond)

- (i) (a) Keto group (Carbonyl group)  $C=O$   
(b) Amino group (NH<sub>2</sub>)  
(c) Aromatic group (phenyl group)

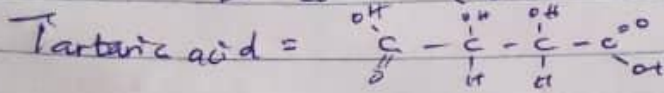
- (ii) (a) Aldehyde group  
(b) Hydroxyl group  
(c) double bond (Alkene group)



No geometric isomer

$$\text{Concentration (mol dm}^{-3}\text{)} = \frac{\text{Conc (g dm}^{-3}\text{)}}{\text{molar mass (g/mol)}}$$

$$[\alpha]_D^{25} = \frac{\alpha}{c \cdot l}$$



$$\text{molar mass} = 150 \text{ g/mol}$$

$$0.856 \text{ g} - 100 \text{ cm}^3$$

$$2 \text{ g} \rightarrow 100 \text{ cm}^3$$

$$0.856 \times 1000 = 85.6 \text{ g/dm}^3$$

L $\phi$

$$\text{Concentration in g/cm}^3 = \frac{\text{Concentration (g/dm}^3\text{)}}{1000}$$

$$= \frac{85.6}{1000} = 0.0856 \text{ g/cm}^3$$

$$[\alpha]_D^{25} = \frac{\alpha}{c \cdot l} = \frac{+1.0}{0.10 \times 50} \neq 11$$

$$[\alpha]_D^{25} = \frac{\alpha}{c \cdot l}; \alpha = +1.0 \quad c = \frac{0.856}{10} = 0.0856 \text{ g/cm}^3$$

$$= \frac{+1.0}{0.0856} = 11.67^\circ$$