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Mat no: 17/MHS01/138
Det: MBBS

1. They have two different ways of classification

a. They are classified based on the number of hydrogen atoms attached to the carbon atom containing the hydroxyl group.

- If it is more than two hydrogen atoms \rightarrow 1° (Primary alcohol)
 - If it is just one hydrogen atom \rightarrow 2° (Secondary alcohol)
 - If it is none \rightarrow 3° (Tertiary alcohol)
- Example: $\text{C}_2\text{H}_5\text{OH}$ (1° primary alcohol)

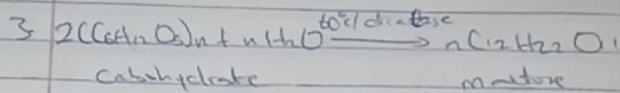
b. They are classified based on the number of hydroxyl groups they possess.

- \rightarrow Monohydric alcohol \rightarrow If there is only one hydroxyl group
- Dihydric alcohol \rightarrow If there are two hydroxyl groups
- Trihydric alcohol \rightarrow If there are three hydroxyl groups
- Polyhydric alcohol \rightarrow If there are more than three hydroxyl groups

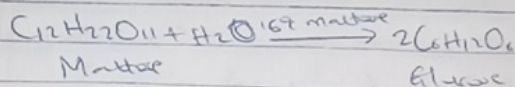
Example: $\text{HOCH}_2\text{CH}_2\text{OH}$ (Ethane-1,2-diol) (dihydric alcohol)

2. For water: All lower alcohols with up to three carbon atoms in their molecules are soluble because they can form hydrogen bonds with water molecules.

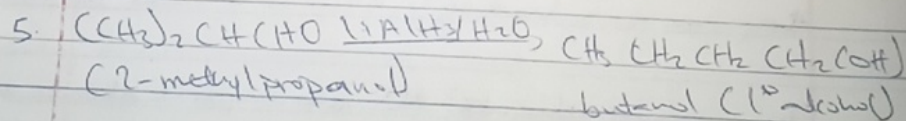
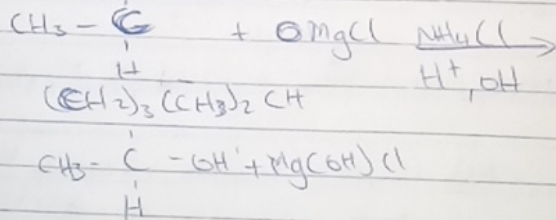
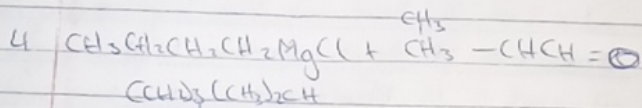
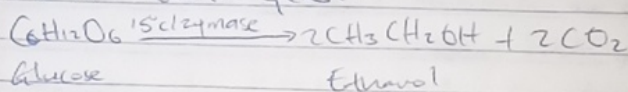
For organic solvents: All monohydric alcohols are soluble in organic solvents.



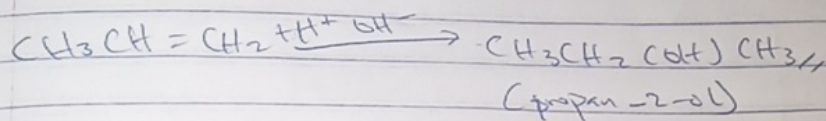
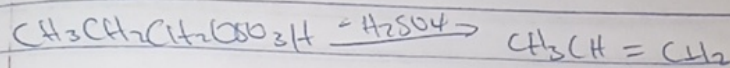
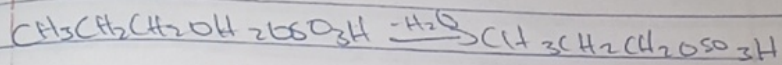
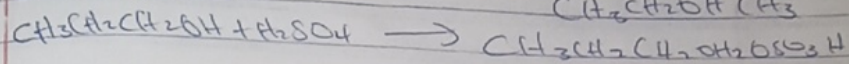
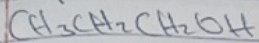
The maltose is broken down into glucose on addition of yeast which contains the enzymes maltase and at a temperature of $15^\circ C$



The glucose at constant temperature of $15^\circ C$ is then converted into alcohol by the enzyme zymase contained also in yeast.



6 Propan-1-ol



Propan-2-ol

