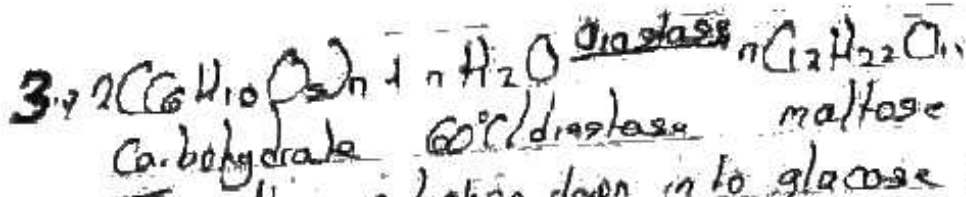


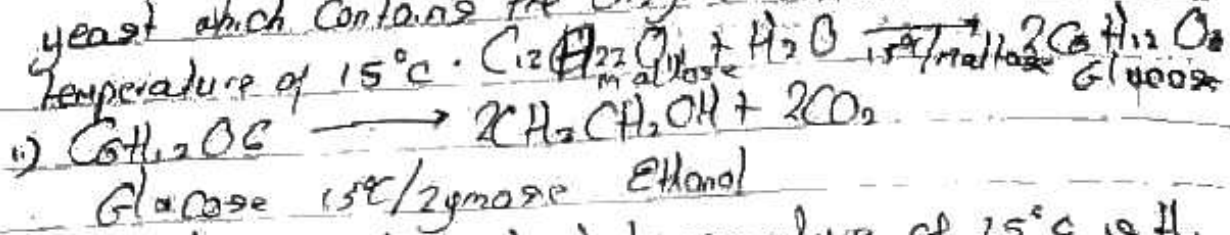
NAME: OMONTUEMHEN PATIENCE NIE
DEPT: NURSING
COLLEGE: MEDICINE AND HEALTH SCIENCES
MATHS No: 19/MHS021099
COURSE No: CHEMISTRY 102

- This is based on the number of hydrogen atoms attached to the carbon atom containing the hydroxyl group. If the numbers of hydrogen atoms attached to the carbon atom bearing the hydroxyl group are three or two, it is called a primary alcohol (1°). If it is one hydrogen atom, it is called secondary alcohol (2°) and if no hydrogen atom is attached to the carbon atom bearing the hydroxyl group, it is called a tertiary alcohol (3°). Example CH_3OH (methanol)
- This is based on the number of hydroxyl groups they possess. Monohydric alcohols have one hydroxyl group present in the alcohol structure. Dihydric alcohols are also called Glycols have two hydroxyl groups present in the alcohol structure while trihydric alcohols or triols have three hydroxyl groups present in the structure of the alcohol. Polyhydric alcohols or polyols have more than three hydroxyl groups. Example $\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$ Propanol (monohydric alcohol)

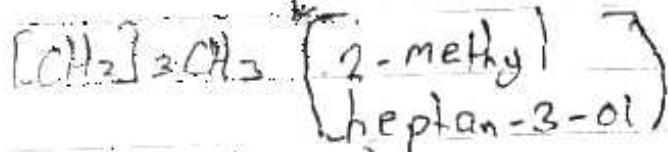
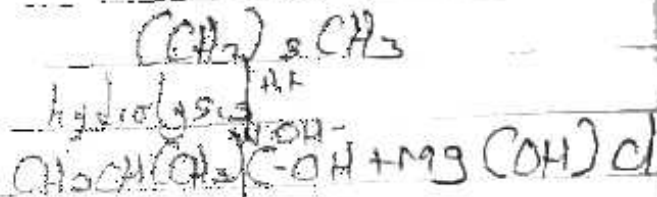
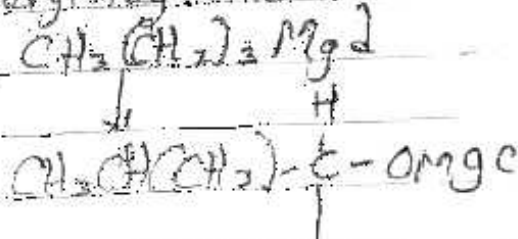
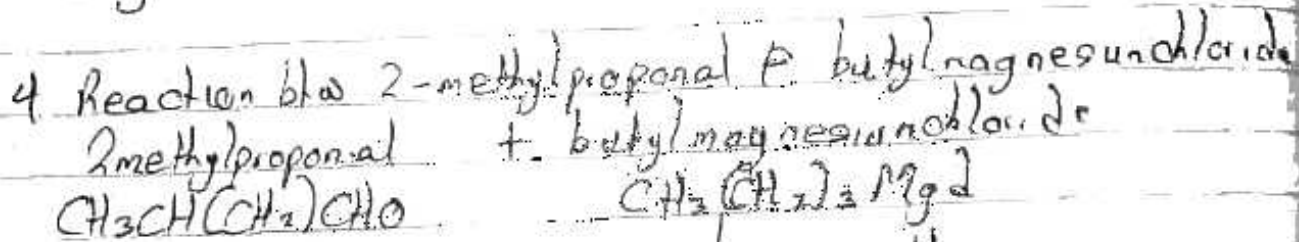
2. Solubility: Lower alcohols with up to three carbon atoms in their molecules are soluble in water, because these lower alcohols can form hydrogen bond with water molecules. The water solubility of alcohols decrease with increasing relative molecular mass. All monohydric alcohols are soluble in organic solvents. The solubility of simple alcohols or polyhydric alcohols is largely due to their ability to form hydrogen bonds with water molecules.



The maltose is broken down into glucose on addition of yeast which contains the enzyme maltase and at a temperature of $15^\circ C$.



The glucose at constant temperature of $15^\circ C$ is then converted into alcohol by the enzyme zymase contained in yeast.

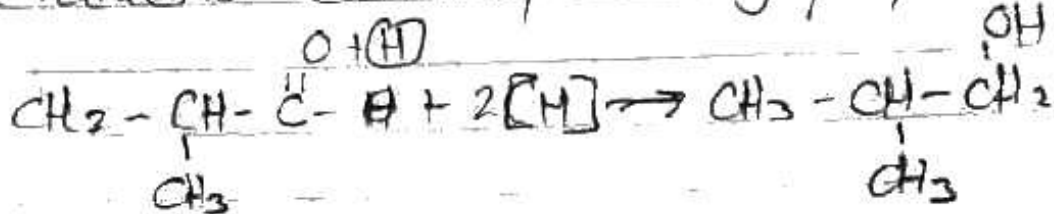


5. ~~Q. 5~~

5. Question was asked to be skipped.

6. Question was asked to be skipped.

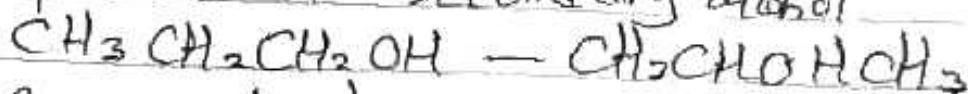
7. The reduction reaction of 2-methylpropanal is



Q. The propose scheme for the conversion of propan-1-ol to propan-2-ol is

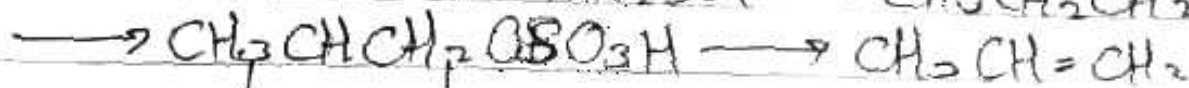
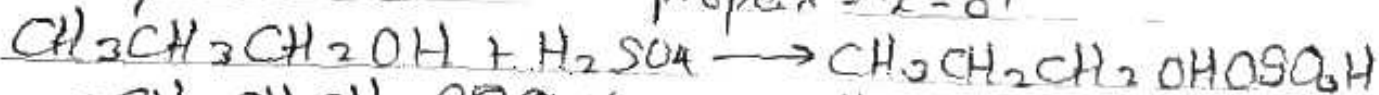
Propan-1-ol - Primary alcohol

Propan-2-ol - Secondary alcohol

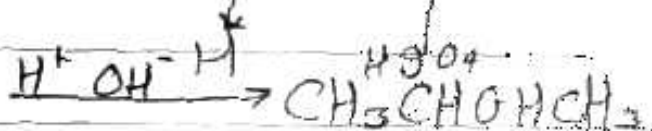


Propan-1-ol

Propan-2-ol



Propene



Propan-2-ol