1. – Methoxyethane

-Ethoxyethane

-Butoxymethane

-Methoxyethane

-Ethoxypropane

2.) Properties of ethers

- Ether molecules cannot form hydrogen bonds with each other

- The lower ethers are highly volatile and flammable

- lower ethers also act as anaesthetics

- the boiling points of ethers and isomeric alcohols becomes lower as the van der Waals forces dominate over the presence of hydrogen bonding.

- Simple ethers are tasteless

3.) Partial dehydration: this is a process that involves excess alcohol with concentrated tetraoxosulphate (IV) acid at the temperature of about 140 degrees Celsius, this is also known as continuous etherification. If excess alcohol is not used then the temperature would have to increase and would yield an alkene instead of an ether.

R–CH2–CH2(OH) → R–CH=CH2 + H2O

Controlled catalytic hydration of olefins:

4.) Uses:

It is used in the hydrolytic production of ethylene glycol

It is used as a gaseous sanitizing agent

It is to produce emulsifying agents, plastics, synthetic textiles etc.