

GT
Ogenji - Gammella Oj

19/ENAG04/039

CH119102

i.) CH_3OCH_3 - Dimethyl ether ~~Methoxyethane~~ Methoxyethane

ii.) $\text{CH}_3\text{CH}_2\text{OCH}_2\text{CH}_3$ - Diethyl ether ~~Ethoxyethane~~ Ethoxyethane

iii.) $\text{X}(\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2)_n\text{O}$ - ~~Pentanamide~~ ~~alkoxypropane~~

iv.) $\text{CH}_3\text{CH}_2\text{OCH}_3$ - methoxyethane

v.) $\text{CH}_3\text{CH}_2\text{CH}_2\text{OCH}_2\text{CH}_3$ - ~~alkoxy~~ ~~alkoxy~~ ~~alkoxy~~

2.) Properties of ether

i.) Physical State: At room temperature ether are colourless, neutral liquid with pleasant odours

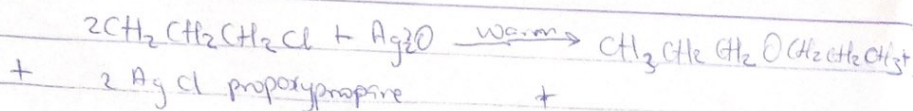
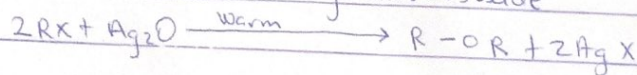
ii.) Density: Most of the simple ethers are less dense than water, although the density increases with increasing relative molecular mass and some of the aromatic ethers are in fact denser than water

iii.) Boiling point: low molecular mass ethers have a lower boiling point than the corresponding alcohols but those ethers containing C_6H_5 groups larger than four carbon atoms, the reverse is true

iv.) Reactivity: Ethers are inert at moderate temperature. Their inertness at moderate temperature leads to their wide use as reaction media. Simple ether are not found commonly in nature but the ether linkage

v.) Solubility: They are less soluble in water than are the corresponding alcohols

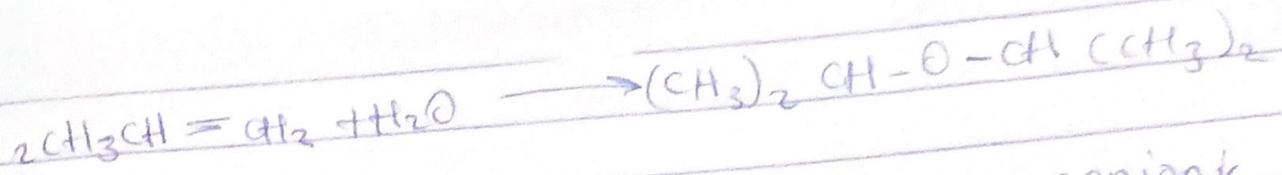
3.) i.) From Haloalkanes and dry silver (I) oxide



ii.) controlled catalytic hydration of olefins



09-21



- 4.) i.) Ethylene oxide is used in the preparation of nonionic emulsifying agent
- ii.) It is used as a gaseous sterilizing agent
- iii.) It is used as an intermediate in the hydrolytic manufacture of ethylene glycol